**DESALINATION BY DIRECTIONAL SOLVENT METHOD**

**RAGHU. K1,DINESH KUMAR. P2and SENTHILKUMAR. S3**

1. Department of Chemistry 2. Department of Civil Engineering 3. Department of Civil Engineering

Krishnasamy College of Engineering & Technology, Cuddalore - 607109, India.

**ABSTRACT**

This project uses the new low temperature, membrane free desalination technology using directional solvents capable of extracting pure water from contaminated solution without themselves dissolving in the recovered water. This method dissolves the water into a directional solvent on increase in temperature, rejects salts and other contaminants, then recovers pure water by cooling back to ambient temperature and reuses the solvent. The directional solvent used here is Decanoic acid (Capric acid). This fatty acid exhibit the required characteristics by having a hydrophilic carboxylic acid end which bonds to water molecules but the hydrophobic chain prevents the dissolution of water soluble salts as well the dissolution of the solvent in water.

**Keywords:** Decanoic acid, Desalination, Salt rejection, Saline water

**I INTRODUCTION**

Most of the desalination processes can be classified as membrane process and evaporative processes. Moreover membrane processes needs high pressure in turn high energy input to overcome osmotic pressure also membranes may foul faster. Other evaporative processes require enormous fuel for heating. The other processes such as Multi Effect Distillation, Multistage Flash distillation etc are under developmental stages (RagavendraPratap Singh, 2013).

A new low temperature, membrane free desalination technology using directional solvents capable of extracting pure water from contaminated solution without themselves dissolves in the recovered water. This method dissolves the water into a directional solvent by increasing its temperature, rejects salts and other contaminants, then recovers pure water by cooling back to ambient temperature and reuses the solvent. The directional solvent used here is decanoic acid. This fatty acid exhibit the required characteristics by having a hydrophilic carboxylic acid end which bonds to water molecules but the hydrophobic chain prevents the dissolution of water soluble salts as well the dissolution of the solvent in water.

By being membrane free process, directional solvent extraction eliminates a source of capital and maintenance cost. Membrane based methods are also unsuitable for highly saline water treatment where high TDS feed water is encountered. By circumventing the use of membranes Directional Solvent extraction promises to become one of the very few currently known techniques to handle a wide TDS range up to and including saturated brines.

Shirui*et al.,* (2018) has emphasized that directional solvent method is less expensive and less energy intensive when compared to the other currently used desalinating techniques, moreover he adds that multi-stage flash uses pressure differences to boil water at a lower temperature than normally required, reverse osmosis uses pressure to recover water from seawater to overcome osmotic pressure and to cross the membrane while directional solvent method uses fatty acids and low-grade waste heat to desalinate. He further explains that ‘Frac’ flowback water can be 3-5 times saltier than seawater and thus much more difficult to treat, but directional solvent method promises to become a feasible way of treating ‘frac’ water, thus it can satisfactorily eliminate the difficulties in desalination (AnuragBajpayee*et al.,* 2011).

**II MATERIALS AND METHODS**

Determination of total suspended solids:A clean porcelain dish is ignited in a muffle furnace and after partial cooling in the air it is cooled in dessicator and weighed.Mix the sample well and pour into funnel with whatman filter paper.A residue sample is placed in the dish and evaporated at1000C on the water bath followed by drying in oven at 1030C for 1 hour.Cool in dessicator and weigh.

Determination of volatile and fixed solids:Initial weight of the crucible (W1), Final weight of the crucible + sample after drying (W2), Final weight of the crucible + sample after drying (W3), Weight of the volatile substance, W= (W2-W3), After determination of Total suspended solids, ignite filter plus crucible at 500C for 20 minutes.

The materials and equipments used in the present work are summarized and provided in the following section under different headings.

1. **Sample collection for seawater:** Seawater was collected from silver beach, Cuddalore.

Silver Beach, which is 57 km long and regarded as one of Asia's longest beaches, is located in Cuddalore, Tamil Nadu, India. It's two kilometers from Cuddalore's downtown on India's southeast coast.

1. **Basic water tests on sea water:** The basic parameters such physical, chemical, biological characteristics of the sample water is tested by pH meter, COD tests, Hardness tests, BOD tests and its range in each parameter is noted.
2. **Desalination process:** Directional solvent method is used for desalinating the sample seawater, the directional solvent used here is Decanoic acid, which is bought from
3. **Extraction of desalinated water:** For heating the directional solvent and extracting desalinated water, hotplate with magnetic stirrer is used.

**Desalination by Directional Solvent Process**

The Directional Solvent Extraction (DSE) method is an innovative desalination technique that utilizes a solvent to extract freshwater from saline water at low temperatures. This process involves heating the directional solvent to dissolve water from the saline solution, then cooling it to precipitate pure water, effectively separating it from salts and other contaminants.

In the DSE method, the directional solvent selectively dissolves water but not salts. When heated, the solvent absorbs water from the saline solution, creating a mixture. Upon cooling, the water precipitates out, leaving the salts behind. This process allows for the extraction of freshwater without the need for membranes or high temperatures, making it an energy-efficient alternative to traditional desalination methods.

**Basic water tests on recovered water:** The recovered water is tested for the physical, chemical and biological characteristics by the corresponding tests.

**III RESULT & DISCUSSION**

**1. Total Suspended Solids (TSS)**

The initial weight of the crucible (W1) was 0.022 g, the combined weight of the crucible and the sample before drying (W2) was 0.040 g, and the final weight after drying (W3) was 0.023 g. Volume of sample (V) was 20 ml. Based on these measurements, the Total suspended solids (TSS) in the given sample were determined to be 50 mg/L. Total suspended solids present in the given sample was found to be 50 mg/l.

The Central Pollution Control Board (CPCB) of India states that the following are normally the acceptable levels for TSS in wastewater: 100 mg/L for domestic sewage and 100 mg/L for industrial effluents. The sample's total suspended solids (TSS) level of 50 mg/L is well within the 100 mg/L allowable limits for home sewage and industrial effluents established by the Central Pollution Control Board (CPCB) of India.

**2. Total Disolved Solids**

The initial weight of the crucible (W1) was 0.022 g, the combined weight of the crucible and the sample before drying (W2) was 0.040 g, and the final weight after drying (W3) was 0.023 g. Volume of sample (V) was 20 ml. Based on these measurements, the volatile suspended solids (VSS) in the given sample were determined to be 850 mg/L.Volatile suspended solids present in the given sample was found to be 850 mg/l.

The Indian standard permissible limits for TDS in various types of water can vary based on the context (e.g., treated effluent, industrial effluents, etc.). However, for typical effluent discharge standards, the permissible limit for TDS is generally around 200 mg/L for most discharge applications as per the Central Pollution Control Board (CPCB) guidelines.

After treatment , The initial weight of the crucible (W1) was 0.022 g, the combined weight of the crucible and sample before drying (W2) was 0.028 g, and the final weight after drying (W3) was 0.024 g. The volume of the sample (V) was 20 mL. Based on these measurements, the volatile suspended solids (VSS) in the sample were calculated to be 200 mg/L. According to the Bureau of Indian Standards (BIS), the upper limit of TDS levels in water is 500 mg/L. The TDS level recommended by WHO, however, is 300 mg/L.

**2. Chemical Oxygen Demand(COD)**

The kind of discharge determines the acceptable COD levels in treated wastewater according to Indian guidelines. (Central Pollution Control Board).   
COD should not be more than 250 mg/L when disposed of in inland surface waters.   
COD should not be more than 100 mg/L when disposed of on land used for irrigation.   
While COD is not typically included in primary drinking water standards, a high COD indicates contamination and unsuitability for drinking water without extensive treatment.

**Before Treatment :** The determination of the oxygen consumed by the wastewater sample was carried out using a burette filled with the standard FAS solution, with 1-2 drops of ferroin indicator as the endpoint detector. The endpoint was noted when the solution changed from blue-green to reddish-brown. The back titre value (V2) was 7, and the blank titre value (V1) was 98. The amount of oxygen consumed by 1000 ml of wastewater was calculated to be 371.28 PPM.

**After Treatment:** The determination of the oxygen consumed by the wastewater sample was carried out using a burette filled with the standard FAS solution, with 1-2 drops of ferroin indicator as the endpoint detector. The endpoint was noted when the solution changed from blue-green to reddish-brown. The back titre value (V2) was 15, and the blank titre value (V1) was 20. The amount of oxygen consumed by 1000 ml of wastewater was calculated to be 204 PPM.

**3. Hardness Test**

**Table No: 1 Titration I - Standardisation of EDTA**

**(Std. Hard Water Vs EDTA)**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **S.No** | **Vol. of hard water(ml)** | **Burette reading** | | **Vol. of EDTA(ml)** | **Concordant value(ml)** | **Indicator** |
| **Initial** | **Final** |
| 1. | 20 | 0 | 19.8 | 19.8 | 19.8 | EBT |
| 2. | 20 | 0 | 19.8 | 19.8 |

The experiment was performed to standardize EDTA using standard hard water. Initially, 20 mL of standard hard water was pipetted into a conical flask. A test tube of ammoniacal buffer solution was added to the flask, followed by a pinch of Eriochrome Black T (EBT) indicator, which resulted in a wine-red color. EDTA solution was then titrated from the burette until a steel blue endpoint was observed. The volume of EDTA required to reach the endpoint was recorded as 19.8 ml.

The standardization of EDTA was carried out using 20 ml of standard hard water, which contains 20 mg of CaCO3. The volume of EDTA consumed, V1​ corresponds to 20 mg of CaCO3, establishing that 1 ml of EDTA equals 0.99 mg of CaCO3.

**Table No: 2.1 Titration II - Estimation of Total Hardness of water  
(Sample water vs EDTA)**

**Before Treatment**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **S.No** | **Vol. of hard water(ml)** | **Burette reading** | | **Vol. of EDTA(ml)** | **Concordant value(ml)** | **Indicator** |
| **Initial** | **Final** |
| 1. | 20 | 0 | 12.7 | 12.7 | 12.7 | EBT |
| 2. | 20 | 0 | 12.7 | 12.7 |

Titration was performed to estimate the total hardness of the given water sample using EDTA as the titrant and Eriochrome Black T (EBT) as the indicator. The volume of hard water taken was 20 mL, and the concordant value of EDTA required for titration was 12.7 ml.Therefore, the total hardness of water sample is 628.65 PPM.

**Table No: 2.2 Titration II - Estimation of Total Hardness of water  
(Sample water vs EDTA)**

**After Treatment**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **S.No** | **Vol. of hard water(ml)** | **Burette reading** | | **Vol. of EDTA(ml)** | **Concordant value(ml)** | **Indicator** |
| **Initial** | **Final** |
| 1. | 20 | 0 | 2.5 | 2.5 | 2.5 | EBT |
| 2. | 20 | 0 | 2.5 | 2.5 |

Titration was performed to estimate the total hardness of the given water sample using EDTA as the titrant and Eriochrome Black T (EBT) as the indicator. The volume of hard water taken was 20 mL, and the concordant value of EDTA required for titration was 2.5 ml.Therefore, the total hardness of water sample is 123.75 PPM.

The total hardness of the water sample is 123.75 ppm, which is acceptable as per Indian Standards (IS 10500:2012), as the permissible limit for total hardness is 200 ppm, with an extended acceptable limit of up to 600 ppm in the absence of an alternative source.

**Table No: 3.1 Titration III -Estimation of Permanent Hardness of Water**

**(Sample water (after boiling) Vs EDTA)**

**Before Treatment**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **S.No** | **Vol. of boiled sample water(ml)** | **Burette reading** | | **Vol. of EDTA(ml)** | **Concordant value(ml)** | **Indicator** |
| **Initial** | **Final** |
| 1. | 20 | 0 | 10 | 10 | 10 | EBT |
| 2. | 20 | 0 | 10 | 10 |

Table No. 3.1 presents the results of the titration to estimate the permanent hardness of water, comparing sample water (after boiling) with EDTA.

For the first titration, 20 ml of boiled sample water was used. The initial burette reading was 0 ml, and the final reading was 10 ml, resulting in a volume of 10 ml of EDTA used, with a concordant value of 10 ml. The indicator used was EBT.

In the second titration, a similar setup was used, with 20 ml of boiled sample water and a burette reading from 0 to 10 ml, resulting in 10 ml of EDTA, though no concordant value was indicated in this case. To express the hardness in PPM, multiply the molar concentration by the molar mass of CaCO₃. Therefore, the permanent hardness of water sample is approximately 495 PPM.

**Table No: 3.2 Titration III -Estimation of Permanent Hardness of Water**

**(Sample water (after boiling) Vs EDTA)**

**After Treatment**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| **S.No** | **Vol. of boiled sample water(ml)** | **Burette reading** | | **Vol. of EDTA(ml)** | **Concordant value(ml)** | **Indicator** |
| **Initial** | **Final** |
| 1. | 20 | 0 | 1.4 | 1.4 | 1.4 | EBT |
| 2. | 20 | 0 | 1.4 | 1.4 |

The permanent hardness of water sample is approximately 69.3 PPM. The Indian Standard for Permanent Hardness in drinking water specifies a desirable level of 200 PPM and a maximum permissible level of 500 PPM. A sample value of 5.5 mg/L indicates very low hardness, signifying soft water that is well within the safe limits for consumption in India."

**Estimation of temporary hardness of water**

**Before treatment:**

The water's permanent hardness is 496 ppm, while the total hardness is 628.65 ppm. By subtracting the permanent hardness from the total hardness, the temporary hardness is calculated to be 132.65 ppm. Temporary hardness is primarily caused by the presence of bicarbonate salts, which can be removed by boiling the water.

**After treatment:**

The water's permanent hardness is 69.3 ppm, while the total hardness is 123.75 ppm ppm. By subtracting the permanent hardness from the total hardness, the temporary hardness is calculated to be 54.45 ppm. Temporary hardness is primarily caused by the presence of bicarbonate salts, which can be removed by boiling the water.

**Result:**

In this project, we had demonstrated Directional Solvent Extraction (DSE) as a new membrane free, low temperature method for desalination of water with applications in production of drinking water from the oceans as well in treatment of flow back and produced water from unconventional oil and gas extraction. DSE is demonstrated to deliver >97% salt rejection for samples of salinity similar to seawater and >95%for salinity similar to produced water. For seawater concentrations DSE is demonstrated to achieve recovery ratios of over 70% with a single stage process and almost 85%with a three-stage process. It is shown that DSE could operate at low temperatures and therefore may become an effective way of utilizing abundant and inexpensive waste heat. In addition, it is shown that DSE has the potential of being effective and economical as a process for desalination of high TDS waters.

Regarding the seawater sample from Silver Beach, Cuddalore, the provided after-treatment values indicate the following:

**Table No: 4 Comparision of Before and After Treatment**

|  |  |  |  |
| --- | --- | --- | --- |
| **S.No** | **Particulars** | **Before Treatment** | **After Treatment** |
| **1.** | **Total suspended solids (TSS)** | 50 mg/L | - |
| **2.** | **Total Disovled solids (TDS)** | 850 mg/L | 200mg/l |
| **3.** | **Chemical Oxygen Demand (COD)** | 371.28 PPM | 204 PPM |
| **4.** | **Total hardness** | 628.65 ppm | 123.75 ppm |
| **5.** | **Permanent hardness** | 496 ppm, | 69.3 ppm |
| **6.** | **Temporary hardness** | 132.65 ppm | 54.45 ppm |

These values suggest that the seawater has undergone some form of treatment, as typical seawater has higher turbidity, TDS, and hardness levels. For instance, untreated seawater generally has a TDS of about 850 mg/L, which is significantly higher than the 200 mg/L reported. While specific pre-treatment values for the DSE method are not provided in the available literature, it is known that the effectiveness of the DSE method can be influenced by factors such as raw water quality, climatic conditions, and seasonal variations. In summary, the DSE method offers a promising approach to desalination by utilizing a solvent to extract freshwater from saline water at low temperatures.

**Volatile Suspended Solids:** The water sample's volatile suspended solids (VSS) content was 200 mg/L, much below the allowable limit of less than 500 mg/L for drinking water. This suggests that the water sample is appropriate for consumption or treatment procedures due to its high solids concentration.  
**Chemical Oxygen Demand (COD):** The water sample's COD of 204 PPM was nearer to the 200 mg/L maximum that is strictly allowed for discharge on land for irrigation. Since there is little organic contamination indicated by this low COD, the wastewater is probably clean and doesn't need to be thoroughly treated before being disposed of.  
**Total Hardness:** According to Indian guidelines, the sample's total hardness of 123.75 PPM is significantly lower than the 200 PPM maximum allowed for drinking water. This suggests that there is no chance of scale building in plumbing systems due to the water's softness.  
The sample's permanent hardness of 69.3 PPM, which is likewise within the safe drinking water standards, indicates that it is soft water and may be consumed without worrying about hardness-related problems. In conclusion, the water sample is safe for consumption, discharge into the environment, or use in industry because it satisfies the requirements for drinking water and has low levels of VSS, COD, and hardness.

**Reference**

1. AnuragBajpayee, T. Luo, A. J. Muto, C. Chen,( 2011), Very low temperature free desalination by directional solvent extraction, *Energy and Environmental Science Journal*, 4, ISSN:1672-1675.
2. RagavendraPratap Singh, (2013), Water Desalination “The Role of RO and MSF”, 2013*, IOSR Journal of Environmental Science and Toxicology and Food Technology*, ISSN:2319-2399,Vol.6. pp. 61-65.
3. Rummi Devi Saini, (2017), Health Risks from Long Term Consumption of Reverse Osmosis Water*, International Journal of Applied Chemistry,* ISSN:0973-1792, Vol.13, pp.293-301.
4. ShiruiLuo, Yunsong Pangand TengfeiLuo,(2018), A Continuous Directional Solvent Extraction Desalination Process Realized with the Aid of Electro-coalescence,*Journal of Chemical Engineering & Process Technology*, ISSN: 2157-7048, Vol.9.1-7.
5. ShiruiLuo, Yunsong Pangand TengfeiLuo,(2018), A Continuous Directional Solvent Extraction Desalination Process Realized with the Aid of Electro-coalescence,*Journal of Chemical Engineering & Process Technology*, ISSN: 2157-7048, Vol.9.